

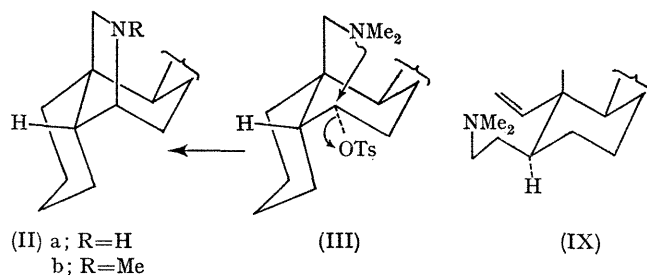
## Novel Heterocyclic Steroidal Amines

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**Summary** Interrelated photochemical and thermal reaction sequences are outlined with novel stereochemical or other mechanistic features leading to derivatives of 6 $\beta$ ,19-imino-5 $\alpha$ - and -5 $\beta$ -cholestanes (I and II) and a range of aza-steroids such as (VIII) containing 2-methylpiperidine residues.

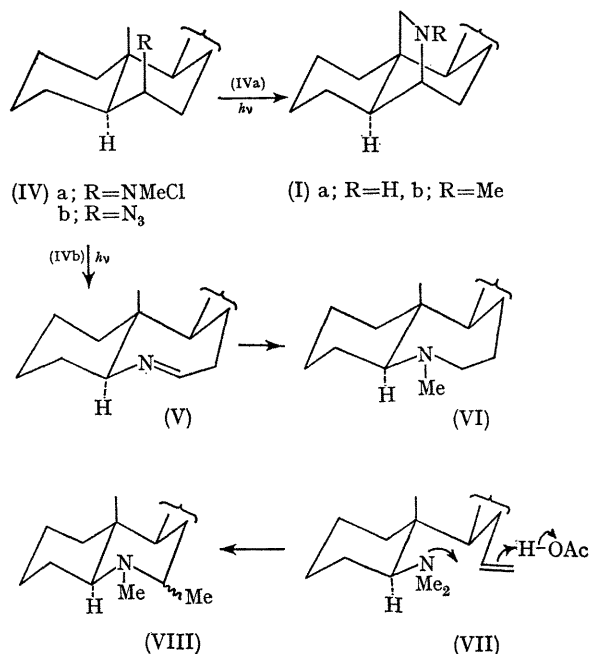
We described previously<sup>1</sup> the cyclisation of a steroidal 19-dimethylamino-6 $\beta$ -tosylate to the methotosylate of a bridged heterocyclic system formulated as 6 $\beta$ ,19-methyl-imino-5 $\alpha$ -cholestane (Ib). The precursor 6 $\alpha$ -alcohol had been prepared from the corresponding 5 $\alpha$ -ketone by reduction with sodium and pentanol. It now appears that the equilibrium in alkaline solution between the 5 $\alpha$ - and 5 $\beta$ -ketones may be markedly affected by the angular dimethylamino-methyl group, with the subsequent result that the tosylate cyclised actually had the 5 $\beta$ -configuration (III) leading to the methotosylate of the amine (IIb). However, the crystalline ketone, m.p. 59°, previously<sup>1</sup> described was apparently correctly formulated as the 5 $\alpha$ -ketone, a balance of conformational and crystal-lattice forces being operative in the synthetical sequence. The amplitude of the o.r.d. Cotton-effect curve of a solution of the 5 $\alpha$ -ketone in 0.5-N-alcoholic KOH increases nearly threefold in 24 hr. at room temperature, suggesting<sup>2</sup> substantial conversion into the 5 $\beta$ -isomer.



An unambiguous synthesis of the base (I; R = Me) has now been effected by photochemical cyclisation of the chloramine (IVa), previous difficulties<sup>1</sup> in handling this very unstable intermediate being successfully overcome by its isolation from solution in ether-cyclohexane by freeze-drying rather than evaporation, followed by photolysis at *ca.* 10° in trifluoroacetic acid. The methiodides from (Ib) and (IIb) have, respectively, m.p. 142–143°, N-CH<sub>3</sub> at  $\tau$  6.08 and 6.30; and m.p. 250–251°, N-CH<sub>3</sub> at  $\tau$  6.22 and 6.36 (n.m.r. spectra in CDCl<sub>3</sub> solutions).

Apart from 6-imino-5 $\alpha$ -cholestane (*ca.* 40%) the chief nitrogenous product (*ca.* 30%) from the photolysis of 6 $\beta$ -azido-5 $\alpha$ -cholestane (IVb) in refluxing cyclohexane appears to be the aza-B-homocholestene (V) rather than the initially

expected<sup>3</sup> heterocycle (Ia). Reductive methylation of (V) with formaldehyde and formic acid gives 6-methyl-6-aza-B-homo-5 $\alpha$ -cholestane (VI) identical with an authentic specimen prepared from 5 $\alpha$ -cholestan-6-one oxime by successive Beckmann rearrangement, reduction of the resultant lactam with lithium aluminium hydride, and *N*-methylation.



A group of related azahomocholestanes including (VI) has been converted into azacholestanes [*e.g.*, (VIII), as methoacetate] by Hofmann degradation followed by subsequent cyclisation of the methines by heating (160 hr., reflux) in acetic acid<sup>4,5</sup> [*cf.* (VII)]. Of the range of methines examined (from syntheses starting with 5 $\alpha$ -cholestan-3-, -4-, -6-, and -7-one oximes), only (IX) from 3-methyl-3-aza-A-homo-5 $\alpha$ -cholestane failed to cyclise in acetic acid, but others did so even in hot ethylene glycol.<sup>5</sup> Such cyclisations in acetic or weaker acids are probably the synchronous processes represented, in contrast to the reactions with hydrogen bromide where the intermediate  $\epsilon$ -bromoamine salt may be isolated [and sometimes even the free base, as from the methine (IX)]. The detailed mechanism of these reversed Hofmann eliminations is being studied.

(Received, August 1st, 1969; Com. 1177.)

<sup>1</sup> R. Ledger and J. McKenna, *Chem. and Ind.*, 1963, 1662.

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<sup>3</sup> D. H. R. Barton and L. R. Morgan, jun., *J. Chem. Soc.*, 1962, 622; *cf.* D. H. R. Barton and A. N. Starratt, *ibid.*, 1965, 2444.

<sup>4</sup> J. McKenna and A. Tulley, *J. Chem. Soc.*, 1960, 945; K. Jewers and J. McKenna, *ibid.*, p. 1575.

<sup>5</sup> *cf.* H. Favre, R. D. Haworth, J. McKenna, R. G. Powell, and G. H. Whitfield, *J. Chem. Soc.*, 1953, 1115.